

Chapter 15 Water And Aqueous Systems Vocabulary Review

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CHAPTER 15 | Aqueous Equilibria: Chemistry of the Water World 15.1. Collect and Organize Figure P15 shows .1 four lines to describe the possible dependence of

percent ionization of acetic acid with concentration. We are to choose the one that best represents the trend for this weak acid. Analyze

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SECTION 15.1 WATER AND ITS PROPERTIES (pages 445-449)

Chapter 15 Water and Aqueous Systems. Chapter 15 “Water and Aqueous Systems”. The Water Molecule: a Review. • Water is a simple tri-atomic molecule, H. 2. O. • Each O-H bond is highly polar, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values) • bond angle of water = 105°.

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Chapter 15 - Water and Aqueous Systems - 15.1 Water and Its Properties - 15.1 Lesson Check - Page 493: 2 Answer The structure of ice is a regular open framework of water molecules arranged like a honeycomb.

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Chapter 15 — Water & Aqueous System. the chaotic movement of colloidal particles, caused by collision with particles of the solvent in which they are dispersed. a compound that conducts an electric current when it is in an aqueous solution or in the molten state; all ionic compounds are electrolytes, but most covalent compounds are not.

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Chapter 15 Water and Aqueous Systems 159 SECTION 15.1 WATER AND ITS PROPERTIES (pages 445–449) This section describes the properties of water in the liquid and solid states and explains how hydrogen bonding affects the surface tension and vapor pressure of water. Water in the Liquid State (pages 445–447) 1.

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